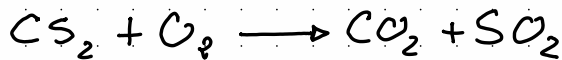
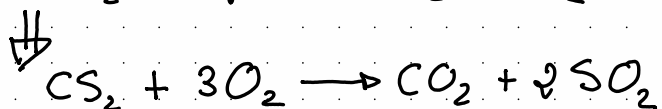


1A



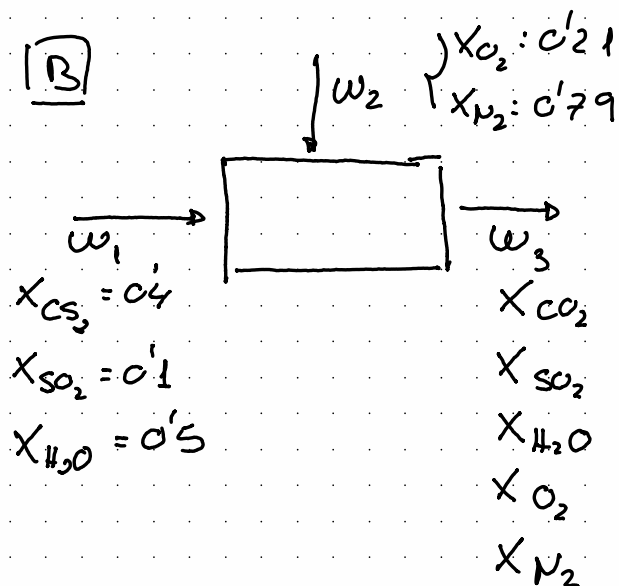
$\Downarrow$



SO_2 po se oxida mais

H_2O po se oxida mais

1B



1C Base de cálculo $\Rightarrow w_1 = 100 \text{ mol/s}$

1D $E + G = S + C$

$$E = 10 \text{ mol/s}$$

$$G =$$

$$40 \frac{\text{mol } CS_2}{1} \times \frac{2 \text{ mol } SO_2}{1 \text{ mol } CS_2} = 80 \frac{\text{mol } SO_2}{1}$$

Summary of values:

- $S = 90$
- $\frac{\text{mol } SO_2}{1}$
- $S = ?$
- $C = \phi$

E

$$E + G = S + C$$

$$E = \emptyset$$

$$G =$$

$$40 \frac{\text{mol CS}_2}{1} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol CS}_2} = 40 \frac{\text{mol CO}_2}{1}$$

$$S = ?$$

$$C = 0$$

$$\left. \begin{array}{l} S = 40 \\ \text{mol CO}_2 \\ 1 \end{array} \right\}$$

F

$$E + G = S + C$$

$$E = ?$$

$$G = \emptyset$$

$$S = ?$$

$$C =$$

$$40 \frac{\text{mol CS}_2}{1} \times \frac{3 \text{ mol CO}_2}{1 \text{ mol CS}_2} = 120 \frac{\text{mol CO}_2}{1}$$

G

$$E = 120 + 120 \cdot X$$

$$S = 120 \cdot X$$

H

$$(120 + 120X) \frac{\text{mol CO}_2}{1} \times \frac{79 \text{ mol N}_2}{21 \text{ mol CO}_2} =$$

$$= (120 + 120X) \cdot 3.76$$

[I]

$$\frac{SO_2}{CO_2 + SO_2 + \underset{\substack{\downarrow \\ \text{exceso}}}{O_2} + N_2} \leq 0.02$$

¿por qué no se incluye el agua?

[5]

$$\frac{90}{40 + 90 + 120x + (120 + 120x) \cdot 3.76} = 0.02$$

[12]

$x \geq \Rightarrow$ Resuelve e indica el %
exceso de aire